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Chemistry- Chapter 15 Solutions. when a nonvolatile solute is added to a solvent, the solvent becomes

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involved in the  
solvation process.

Fewer solvent particles  
closest to the surface  
are able to escape into  
vapors because they  
are too busy dissolving  
the solute. separated  
easily by filtration or  
settling.

## **Chemistry- Chapter 15 Solutions**

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Chapter 15 - Solutions  
15.1 Definitions related to Solutions  
Solution - uniform mixture of two or more substances. A solution is composed of a solute(s) dissolved in a solvent. Solute - the substance present in lesser amount in a solution.

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Chapter 15 Solutions.

1. A homogeneous

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mixture is a combination of two (or more) pure substances that is uniform in composition and appearance throughout. Examples of homogeneous mixtures in the real world include rubbing alcohol (70% isopropyl alcohol, 30% water) and gasoline (a mixture of hydrocarbons). 2.

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- **Francis Howell**  
**High School**

Identify the solute and the solvent in....

Identify the solute and the solvent in.... Water is a polar solvent;  $\text{CCl}_4$  is a non.... Sodium chloride is the solute, water is the solvent.

Water is the solute, ethanol is the solvent.

Oxygen is the solute, nitrogen is the solvent.

$\text{CCl}_4$ . Identify the solute and the solvent in....

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Kinetics. 2 ... c See

Chemistry Now ,  
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Solution 2 hr and 20  
min = 4 half-lives Half-  
life Time Elapsed Mass  
Left 1st 35 min 2.50 g  
2nd 70 1.25 g 3rd 105  
0.625 g 4th 140 0.313  
g Rate =  $k[\text{sugar}]$  and

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$k = 3.3 \times 10^{-4} \text{ sec}^{-1}$ . Half-life.

Chemistry  
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**Reactivity: Chemical**  
**Kinetics**

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Why do two charge  
bodies attract or repel  
? [NCT2008] Answer:  
When the two bodies  
have the same

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charges, they repel each other and when the two bodies have different charges they attract each other.

Question 2.

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polymers, names and

structure of monomers.

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life covers questions

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Chapter Exercises. 2.  
0.014 M. 4.  $7.2 \times 10$   
 $-15$  M. 6.  $4.4 \times 10^{-22}$   
M. 8.  $6.2 \times 10^{-6}$  M =  
[Hg  $2+$ ];  $1.2 \times 10^{-5}$  M

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= [Cl<sup>-</sup>]; The substance is a weak electrolyte because very little of the initial 0.015 M HgCl<sub>2</sub> dissolved. 10. [OH<sup>-</sup>] =  $4.5 \times 10^{-5}$ ; [Al<sup>3+</sup>] =  $2.2 \times 10^{-20}$  (molar solubility) 12.

## 15.3 Multiple Equilibria - Chemistry

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## Hydrocarbons

(compounds composed of carbon and

hydrogen) in which all of the carbon-carbon bonds are single bonds are called alkanes. 10.

When a(n) small section of an organic molecule is largely responsible for the molecule's chemical and physical characteristics, that section is called a functional group.

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15. \_\_\_\_ The elixir of  
life was supposed to  
convey immortality to  
humans. 16. \_\_\_\_ Pure  
chemistry always has a  
practical goal. 17. \_\_\_\_

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The development of wrinkle-free fabrics is an example of applied chemistry. 18. \_\_\_\_\_

Studying how fast crystals form from salt solutions would be done by a physical chemist. 19. \_\_\_\_\_

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